

Subject card

Subject name and code	Analytical chemistry, PG_00052416						
Field of study	Chemical Business						
Date of commencement of studies	October 2024	Academic year of realisation of subject			2025/2026		
Education level	Bachelor's studies	Subject group			Obligatory subject group in the field of study		
Mode of study	full-time studies	Mode of delivery			at the university		
Year of study	2	Language of instruction			Polish		
Semester of study	3	ECTS credits			2.0		
Learning profile	academic	Assessment form			exam		
Conducting unit	Department of Analytical Chemistry -> Faculty of Chemistry -> Rector						
Name and surname of lecturer (lecturers)	Subject supervisor		dr Dorota Zarzeczańska				
	Teachers		dr Dorota Zarzeczańska				
Lesson types	Lesson type	Lecture	Tutorial	Laboratory	Project	Seminar	SUM
	Number of study hours	30.0	0.0	0.0	0.0	0.0	30
	E-learning hours included: 0.0						
Learning activity and number of study hours	Learning activity	Participation in didactic classes included in study plan		Participation in consultation hours		Self-study	SUM
	Number of study hours	30		5.0		15.0	50
Subject objectives	<p>introduction of the principles of division of cations and anions into analytical groups,</p> <p>acquainting with the basic methods used in the quantitative and qualitative analysis of inorganic compounds,</p> <p>introduction of the methods of identifying cations and anions</p>						

Learning outcomes	Course outcome	Subject outcome	Method of verification
	[BCHINŻ_K04] Demonstrates responsibility for the safety of her/his own and others' work.	Recognizes and predicts sources of errors in analysis and follows occupational health and safety rules in the laboratory.	[SK1] oral statement/conversation/discussion
	[BCHINŻ_U03] Plans, selects the appropriate research and measuring equipment and performs simple chemical experiments; analyses the results and draws conclusions based on them.	plans and selects laboratory equipment and glassware necessary to perform classical analysis	[SU4] test/exam - oral or written
	[BCHINŻ_W10] Applies safety and hygiene principles when working on a test and measurement stand or in the field.	Characterizes the basic principles of health and safety procedures at the analytical laboratory.	[SW4] test/exam - oral or written
	[BCHINŻ_W02] Enumerates basic laws and theories in chemistry, physics and mathematics necessary to formulate and solve simple engineering tasks.	<ol style="list-style-type: none"> 1. Provides the composition of group reagents. 2. Explains the working principles of group reagents and analyte-specific reagents. 3. Defines the basic problems of the theory describing the course of ionic reactions in solution. 4. Illustrates the course of the titration with the appropriate curve. 5. Illustrates and describes by means of chemical equations reactions taking place during qualitative and quantitative determinations. 6. Lists and explains the modus-operandi of indicators used in the quantitative titration. 	[SW4] test/exam - oral or written
	[BCHINŻ_U08] Uses the chemical nomenclature and engineering terminology properly.	Appropriately uses chemical nomenclature and engineering terminology typical of analytical chemistry.	[SU4] test/exam - oral or written
	[BCHINŻ_W07] Describes the construction and operating principles of basic scientific, technological and control-measuring apparatus.	describes and can sketch laboratory equipment and glassware necessary to perform classical analysis	[SW4] test/exam - oral or written
Subject contents	Chemical reactions in analytical chemistry. Equilibria in solutions. Determination and detectability of metal ions, anions and inorganic compounds. Collection and preparation of samples for analysis. Basic concepts of classical qualitative analysis. Analytic division of cations by Fresenius. Group reagents and conditions for their use. Characteristic reactions of cations and analytical effects. Division of anions into analytical groups according to Bunsen, characteristic reactions of selected anions. Basic concepts of classical quantitative analysis. Titration analysis - general part, division of titration methods (alkalimetry, redoximetry, complexometry, precipitation titration analyzes), EqP (equivalence point) and EP (end point) concepts, types of titrimetric methods (direct, indirect and inverse). Weight analysis - phenomena related to precipitation and dissolution of sediments. Evaluation of the analysis results.		
Prerequisites and co-requisites	<p>-completed general chemistry course</p> <p>-using basic laboratory glass and applying the rules of work in a chemical laboratory, writing chemical reactions taking into account the stoichiometry of reactions and determining the products, e.g. sediment, gas, etc., describing chemical equilibrium in the solution using chemical reactions, balancing the oxidation and reduction reactions; calculations based on chemical reactions, calculating molar concentrations, percentages, calculating the pH of electrolytes</p>		
Assessment methods and criteria	Subject passing criteria	Passing threshold	Percentage of the final grade
	written exam: accounting tasks (50%), open questions (20%) and closed questions (30%)	51.0%	100.0%

Recommended reading	Basic literature	<p>J. Minczewski i Z. Marczenko, Chemia analityczna 1 i 2, PWN Warszawa</p> <p>T. Lipiec, Z.S. Szmal, Chemia analityczna z elementami analizy instrumentalnej, PZWL Warszawa</p> <p>H. Bentkowska, Chemia analityczna jakościowa, skrypt PG</p> <p>A. Cygański, Chemiczne metody analizy ilościowej, WNT</p> <p>A. Persony , Chemia analityczna. Podstawy klasycznej analizy ilościowej, Medyk</p>
	Supplementary literature	<p>D. Harvey, Modern Analytical Chemistry, McGraw Hill Companies, Inc.</p> <p>W. Gorzelany, A. Śliwa, J. Wojciechowska, Pólmikroanaliza jakościowa, PWN Warszawa</p>
	eResources addresses	

Example issues/
example questions/
tasks being completed

Sample open tasks for the oral and written exam

1. Characterize the conditions for the precipitation of group I cations. List the ions separated. Write reactions with a group reagent. Why can't concentrated hydrochloric acid be used to precipitate chlorides of group I cations?
2. Why do silver chloride and bromide dissolve in ammonia solution, but silver iodide does not dissolve?
3. What is AKT and how does it apply to the analysis of group II cations?
4. Why, when detecting cadmium (II) and copper (II) ions next to each other, during the analysis after dissolving the sulfides in concentrated nitric acid, do we add ammonia to make the reaction alkaline and only then add cyanide ions? Write the reactions described above and describe their color effects. What would happen if we added cyanide ions to an acidic solution?
5. Discuss in detail what group III reagent is. cations and why?
6. What is the systematic analysis of cations according to Fresenius?
7. What is fractional precipitation of halides?
8. What is the division of anions into analytical groups according to Bunsen? Provide example anions for each group.
9. How to detect bromide and iodide ions side by side using chlorine water? Provide the colors of substances present in aqueous and organic solutions during this determination.
10. What is the annular reaction? What ions do we use it to remove? Record the reactions taking place and describe the analytical effects.
11. What are alkalimetric indicators and what is their color change - describe with an example.
12. Draw a titration curve of sodium hydroxide with hydrochloric acid solution. Mark the equivalence point and titration step. Remember to label the axis.
13. What is it and what properties should the basic substance have?
14. Write a redox reaction in which the electrochemical potential depends on pH. Write the Nernst equation for this reaction.
15. Provide the types of indicators used in redoximetry (3) and describe their operation.
16. Discuss setting the titer of sodium thiosulfate (give the basic substance, write the reaction, give the indicator, etc.).
17. Discuss the iodometric determination of copper (substrates, what and why, the sequence of activities performed, indicator, titrant, reactions taking place during the determination).
18. Describe the determination of chloride using Mohr's method. What is the chromate error and how do we prevent it?
19. List and briefly describe precipitation methods for chloride determination.
20. What are metal indicators in complexometry? What conditions should an indicator for complexometric titration meet? Give an example of a metal indicator.

	<p>21. What is EDTA? - provide the structural formula, full name, application, properties.</p> <p>22. Describe the gravimetric determination of iron(III). Why should a precipitate not be left above the solution when determining iron (III) by weight?</p> <p>23. Describe the effect of a foreign ion in the precipitation of sediments, justify your answer?</p> <p>24. Describe the common ion effect in the precipitation of sediments, justify your answer?</p> <p>25. What are the basic types of sediments you know, briefly describe their properties</p> <p>.26. Discuss the types of complexometric determinations. Prześlij opinię Panele boczne Historia Zapisane</p>
Work placement	Not applicable

Document generated electronically. Does not require a seal or signature.