University of Gdańsk

Subject card

Subject name and code	General chemistry, PG_00140696						
Field of study	Natural Resources Conservation						
Date of commencement of studies	October 2024		Academic year of realisation of subject	2024/2025			
Education level	undergraduate studies		Subject group	Obligatory subject group in the field of study			
Mode of study	full-time studies		Mode of delivery	at the university			
Year of study	1		Language of instruction	Polish			
Semester of study	1		ECTS credits	1.0			
Learning profile	academic		Assessment form				
Conducting unit							
Name and surname of lecturer (lecturers)	Subject supervisor	prof. dr hab. Jolanta Kumirska					
	Teachers	dr Małgorzata Czaja prof. dr hab. Jolanta Kumirska					
Lesson types	Lesson type	Lecture	Tutorial	Laboratory	Project	Seminar	SUM
	Number of study hours	0.0	15.0	0.0	0.0	0.0	15
	E-learning hours included: 0.0						
	Additional information: Solving tasks and other simple problems related to the application of acquired knowledge.						
Learning activity and number of study hours	Learning activity	Participation in didactic classes included in study plan	Participation in consultation hours	Self-study	SUM		
	Number of study hours	15	2.0	8.0	25		
Subject objectives	Auditorium exercises: Developing the ability to apply acquired knowledge to solve specific problems and tasks						

Learning outcomes	Course outcome	Subject outcome	Method of verification
	[OZPL3_W09] The graduate possesses an advanced comprehension of the current state of knowledge and the latest trends in biology, as well as their relationship to other natural disciplines	Student defines the most important chemical laws and concepts governing phenomena occurring in nature. Student indicates the relationship between the structure of the atom and the properties of the element and its position in the periodic table. Student lists the most important types of chemical bonds. Student describes the structure of gases, liquids and solids in terms of the kinetic-molecular model of matter. Student defines molar and percentage concentration. Student describes the most important aspects of energy, kinetics and equilibrium of reactions. Student describes the acid-base properties of aqueous solutions using the concept of pH. Student explains the basic concepts of oxidation-reduction reactions and electrochemical phenomena.	[SW1] oral statement/ conversation/discussion [SW3] text preparation/written work
	[OZPL3_W10] The graduate possesses a comprehensive understanding of current issues in biology and related fields	Student describes the basic methods of testing the properties of chemical substances. Student lists the principles of safe handling of hazardous substances. Student lists the most important elements of laboratory equipment and describes their applications for specific activities.	[SW1] oral statement/ conversation/discussion [SW3] text preparation/written work
	[OZPL3_U06] The graduate is able to make observations and perform basic physical, biological and chemical measurements in the field or laboratory	Student organizes and analyzes the results of experiments and draws correct conclusions based on them. Student conducts calculations using the known chemical laws and presents the experimental results in written form.	[SU1] oral statement/conversation/ discussion [SU3] text preparation/written work
	[OZPL3_K03] The graduate is ready to identify priorities to achieve a task defined by him/ herself or others	Student uses acquired knowledge and skills to plan and carry out assigned tasks.	[SK1] oral statement/conversation/ discussion [SK3] text preparation/written work [SK8] observation of student's independent or team work
	[OZPL3_K07] The graduate is prepared to demonstrate responsibility for the equipment/ materials entrusted and respects the work of others	Student demonstrates responsibility for the assigned equipment and reagents, and ensures cleanliness and order at the workplace. Student takes tests and colloquiums and submits reports within the prescribed deadlines.	[SK1] oral statement/conversation/ discussion [SK3] text preparation/written work [SK8] observation of student's independent or team work
	[OZPL3_K06] The graduate is prepared to demonstrate responsibility for their own and others' safe working conditions in the laboratory and in the field, and is able to recognise hazardous situations and take appropriate action	Student follows the rules of safe conduct in a chemical laboratory in such a way as not to pose a threat to one's own health, that of others, or to the environment. Student uses the information contained in the Material Safety Data Sheets.	[SK1] oral statement/conversation/ discussion [SK3] text preparation/written work [SK8] observation of student's independent or team work
	[OZPL3_U01] The graduate is able to use basic apparatus and research tools and maintains the correct sequence of operations in laboratory and field work	Student selects and correctly uses laboratory equipment in accordance with its intended purpose.	[SU1] oral statement/conversation/ discussion [SU3] text preparation/written work
	[OZPL3_U04] The graduate is able to plan and carry out simple research tasks in the biological sciences under the guidance of a supervisor	Student (under the supervision of a supervisor) plans, performs and conducts research experiments.	[SU1] oral statement/conversation/ discussion [SU3] text preparation/written work
Subject contents	Auditorium exercises: Atomic, molecular and molar mass. The relationship of the structure of the atom with its position in the periodic table and the properties of the element. Chemical equation of the reaction, balancing, stoichiometric calculations. Molar and percentage concentration, density of solutions. Calculations related to ionic equilibria in solution: dissociation, hydrolysis, pH.		
Prerequisites and co-requisites	lack		

Assessment methods and criteria	Subject passing criteria	Passing threshold	Percentage of the final grade
	at least 51% in both written tests, positive evaluation of work on student exercises and homework	51.0%	100.0%
Recommended reading	Basic literature	1. Jones L., Atkins P. 2020. Chemia ogólna. PWN, Warsow. 2. Lee J. D. 1994. Zwięzła chemia nieorganiczna. PWN, Warsow. 3. Pauling L., Pauling P. 1997. Chemia. PWN, Warsow.	
	Supplementary literature	1. Bielański A. 2012. Podstawy chemii nieorganicznej. Tom 1, 2. PWN, Warsow.	
	eResources addresses	Adresy na platformie eNauczanie:	
Example issues/ example questions/ tasks being completed			
Work placement	Not applicable		

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